



pH SCALE

The pH scale measures the amount of acidity or basicity in a chemical solution. Each number is ten times larger than the next number. Thus a pH of 1 is 10 times more acidic than a pH of 2; a pH of 2 is 100 times more acidic than a pH of 4. pH is derived from the Greek word for “Potency” and H is the symbol for hydrogen – thus, pH stands for the potency of the hydrogen ion in an aqueous solution.

Note the following:

1. Urine fluctuates from 5.5 to 8.0 depending on age, sex, food consumption, etc.
2. Never mix cleaning chemicals, especially those found on opposite sides of the pH scale because adverse reactions or neutralization of the mixture may occur.
3. All pH values are approximates.
4. Solvents do not have a pH because they do not ionize in aqueous solution.
5. pH does not measure the corrosiveness of an acid (i.e., HCL with a pH of 1 is stronger than Phosphoric with a pH of 1).

